Unit #9 Practice Blackboard Insert

1. **(2 pts.)** What is the percent by mass of a solution which is made by dissolving 43.25 g of $Mg(NO_3)_2$ in 435.0 g of water?

2. **(2 pts.)** How many moles of NH_4NO_3 are present in 336.0 mL of a 0.834 M NH_4NO_3 solution?

3. **(2 pts.)** How many grams of K_2SO_4 would be needed to prepare 715.0 g of a 3.65 % by mass solution of K_2SO_4?

4. **(4 pts.)** Given the following reaction:
   Calcium Nitrate plus Potassium Carbonate Yields Calcium Carbonate and Potassium Nitrate:

   How many milliliters of 0.405 M $Ca(NO_3)_2$ solution are needed to completely react with 33.50 mL of 0.677 M K_2CO_3 solution?

Extra Credit (1 pt):

Convert name to formula or visa versa … from the list of names in unit 5