Acids & Bases

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Acids & Bases: Traditional Properties

Property	Acid	Base
Taste	Sour	Bitter
Feel	None	Slippery
Litmus	B➔R	R➔B
Phenolphthalein	Colorless	Magenta
With Carbonate	CO_2 evolution	None
With "active" Metals	H_2 evolution	None
With most metals	None	Water Insoluble

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Arrhenius Theory: Acids

Acid = substance that forms hydrogen ions in water solution

$$HA_{(aq)} \rightarrow H^+_{(aq)} + A^-_{(aq)}$$

H^+ = proton

But, individual protons do NOT exist in water:

$$H^+ + H_2O \rightarrow H_3O^+ \text{ (Hydronium Ion)}$$

Arrhenius Acids form *hydronium ions* in solution

Arrhenius Theory: Bases

Base = substance that forms hydroxide ions (OH^-) in water

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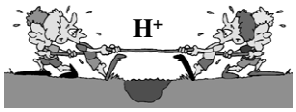
Bronsted-Lowry Theory of Acids & Bases

$$AH + B \rightleftarrows BH^+ + A^-$$

Acid = proton donor
Base = proton acceptor (Prime departure from Arrhenius)



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pH Scale

Measurement of relative acidity

Determined by hydrogen ion concentration

Values range between 0 – 14
- pH < 7 ➔ acidic
- pH = 7 ➔ neutral
- pH > 7 ➔ basic (alkaline)

Measured using

indicators (pH papers or solutions)

pH meter

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pH: A Measure of $[H^+]$ (Molar Concentration of H^+)

$[H^+]$	pH
1×10^{-1}	1
1×10^{-2}	2
1×10^{-3}	3
1×10^{-4}	4
1×10^{-5}	5
1×10^{-6}	6
1×10^{-7}	7
1×10^{-8}	8
1×10^{-9}	9
1×10^{-10}	10
1×10^{-11}	11
1×10^{-12}	12
1×10^{-13}	13
1×10^{-14}	14

$[H^+]$ (Acidity) increasing, pH decreasing

$$[H^+] = 1 \times 10^{-pH}$$

$$pH = -\log [H^+]$$

$[H^+]$ (Acidity) decreasing, pH increasing

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pH Scale

Focus of pH scale is the proton (acidity)

Strong acids: pH < 4 **Strong Bases: > pH 11**

Weak acids: pH 4-6 **Weak Bases: pH 8-11**

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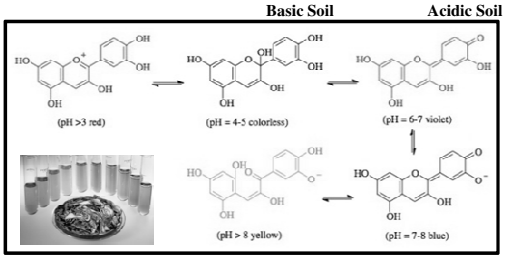
Many Plant Colors (Anthrocyanins) are pH Indicators

Indicators ➔ color depends on pH
Color change ➔ Chemical change

The "colors" in vegetables have significant cancer risk reductions

Hydrangea Basic Soil

Hydrangea Acidic Soil



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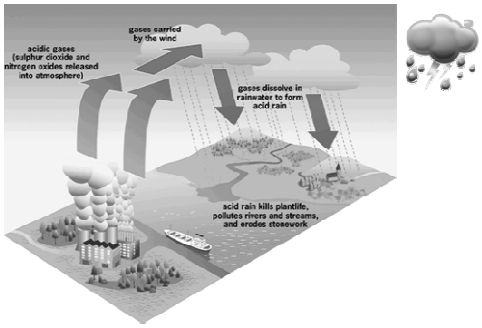
Air-borne Pollution + Water = Acid Rain



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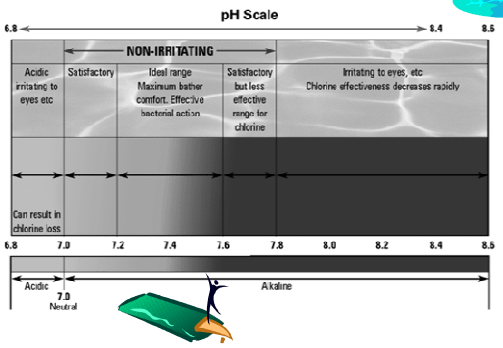
Proper pH Keeps Pools Healthy



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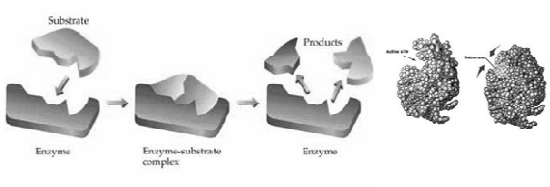
All biochemical reactions have an optimum pH



Changes in pH can change protein shape
&
prevent/alter molecular interactions
Many "genetic diseases" result from incorrect protein shapes

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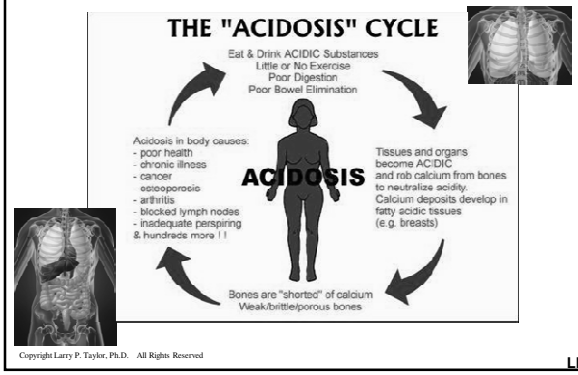
Improper pH Balance Has Many Negative Health Consequences



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Proper pH is important to Plant growth

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Acids & Bases Lab

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Today's Lab (Work in Pairs)

Purpose:

Observe the properties of acids and bases

Use a pH indicator to monitor acidity level

Classify 2 household substances as acids or bases.

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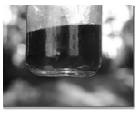
Preparation of Indicator Dye (One Batch per Lab)

Procedure:

- 500 mL of RO water into a 1-L beaker
- Heat on a hot plate
- When the water boils, add ~2 cups of shredded red cabbage
- Boil for 5 minutes
- Filter with a large Buchner funnel into a clean beaker
- Let the purple indicator solution cool while you do part II

Extract



Chemistry department has done this for you!

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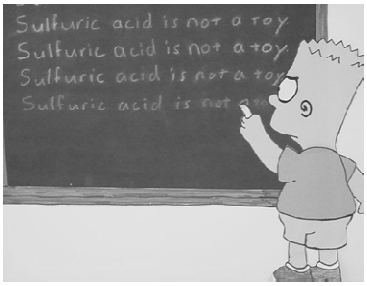
Handle Acids & Bases With Care!



Avoid Contact with acids and bases
Wear your safety goggles
Immediately wash any contact areas with lots of cold water
Notify instructor if you contact any acid or base

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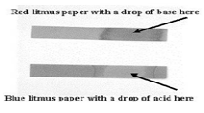
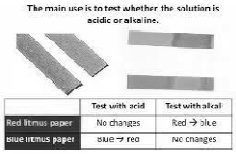
Acids & Bases Properties: Litmus

The acids/bases to be tested: HCl, CH_3COOH, NH_4OH, & NaOH

Litmus Test





	Test with acid	Test with alkali
Red litmus paper	No changes	Red → blue
Blue litmus paper	Blue → red	No changes

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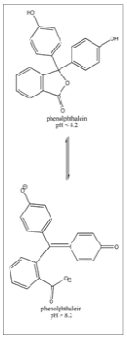
Acids & Bases Properties: Phenolphthalein



One of the most common indicators used

Most common OTC laxative

C.S.I. = used to determine if stain is blood

Kastle-Meyer Spot Test

Phenolphthalein plus sample

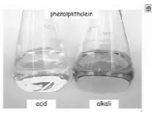
Add H_2O_2

Hemoglobin present oxidizes to pink form

OH^- attacks acid form and changes structure

Acid form: colorless

Basic form: magenta



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Acids & Bases Properties: Metals with Acids

Metals	Metal Ion	Reactivity
K	K^+	reacts with water
Ca	Ca^{2+}	
Na	Na^+	
Mg	Mg^{2+}	reacts with acids
Al	Al^{3-}	
Zn	Zn^{2+}	
Fe	Fe^{2+}	
Ni	Ni^{2+}	
Sn	Sn^{2+}	
Pb	Pb^{2+}	
H_2	H^+	
Cu	Cu^{2+}	highly unreactive
Hg	Hg^{2+}	
Ag	Ag^+	
Pt	Pt^+	
Au	Au^{3+}	

Metals Above Hydrogen

Produce Hydrogen gas

In presence of acid

Mg, Al, Zn, Fe, Ni, Sn, Pb react with acids:

$Zn + H_2SO_4 \rightarrow ZnSO_4 + H_2$

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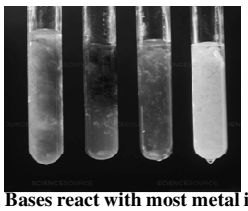
Acids & Bases Properties: Metals with Bases

Metals form insoluble hydroxides



Bases react with most metal ions:

$2\ OH^-_{(aq)} + M^{2+} \rightarrow M(OH)_{2(s)}$

Hydroxide Pollution

Difficult to clean

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Acids & Bases Properties: Carbonates

Acids react with carbonate ions:

$2\ H^+_{(aq)} + CO_3^{-}{}_{(aq)} \rightarrow H_2O_{(l)} + CO_{2(g)}$

Atmospheric $CO_2 + H_2O \rightarrow H_2CO_3$

Dissolves Carbonates
A major erosion process

Geologists test minerals with HCl:
If it "fizzes," it's a carbonate mineral

Carbonates do not react with bases

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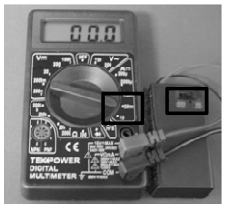
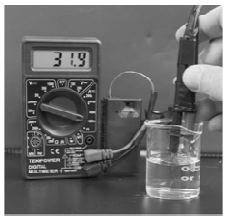
Conductivity





Set meter to 200 m
Set battery to On ⇨
Insert probes
Metal only in solution
Do not touch glass
Read meter
Record value

Turn Battery Off ⇦
Turn Meter Off

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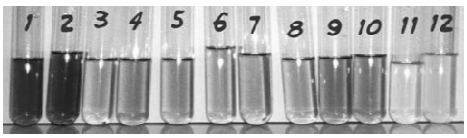
Cabbage Dye Indicator

Use Indicator to infer pH
Test the acidity level of the acids & bases used in Part I & RO water.
Pair A: Test HCl & NH_4OH
Pair B: Test CH_3COOH & NaOH

- **Add two droppers full of the substance to be tested into a small test tube.**
- **Add 3 drops of the cabbage indicator and mix by "flicking" the test**
- **Record the pH of each solution**
- **Combine Results**

pH Indicator Color Chart



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Effect of Concentration on pH (Serial Dilution)

Pair A: Test HCl Pair B: Test NaOH

Label the test tubes 1 and 2.

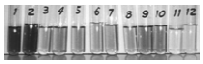
Test tube 1: add 20 drops of acid or base; add 3 drops of the cabbage indicator. Record the color and pH in Table 7.

Dilution of the acid or base:

Measure 100.0 mL of distilled water using the graduated cylinder.
Pour it into the clean 150. mL beaker.
Add one drop of the acid or base to the beaker. Stir with a clean stirring rod.

Test tube # 2: Add 1 mL of the diluted acid or base; add 3 drops of cabbage indicator. Record the color and pH in Table 7.

Exchange data with the other pair in your group to complete Table 7.



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Household Substances

Test 2 Different Household Products (Found in the Hood)

- **Put two droppers full of the substance to be tested into a small test tube**
- **Add 3 drops of the cabbage indicator**
- **Mix well and then record the color and pH of the substance.**
- **Classify each substance as acidic, basic or neutral.**

Data / Observations / Data Interpretation:

Fill in tables

Conclusion

Summarize the characteristic properties of acids and bases (that you observed). Describe the relationship between the pH value and the level of acidity in a solution.

The pH Scale

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Self-Protolysis of Water

$$H_2O + H_2O \rightarrow H_3O^+ + OH^-$$

Case for: ions present; current flows
Case against: no ions present; no current

Typically, H^+ is ~ 10^{-7} (pH 7)
But,
measurement apparatus sensitivity dependent

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Let's Boldly Go Explore Today's Lab

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